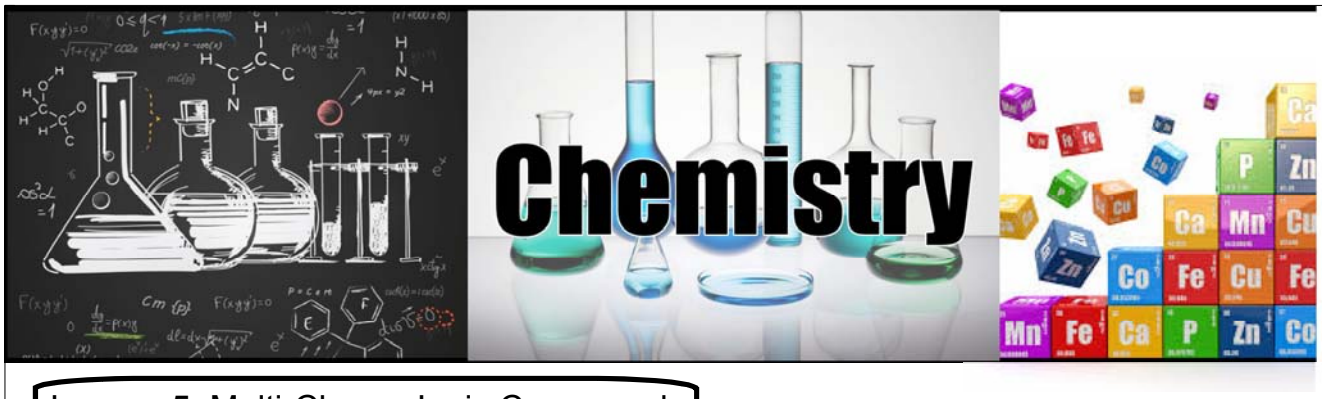

Lesson 5: Multi-Charge Ionic Compounds

- Same as previous lesson.
- Some metals have more than 1 possible charge, the most common is written first
- Example: Fe (iron), Pb (lead), Cu (copper)

iron(II)	2+	2+	1+
iron(III)	3+	4+	2+

- Balance the charges so the compound equals 0.
- Decide which charge to use.

Ex.) Name the following:

a) $PbCl_4$ (4+ 4−) lead(IV) chloride

b) $PbCl_2$ (2+ 2−) lead(II) chloride

c) SnO (2+ 2−) tin(II) oxide

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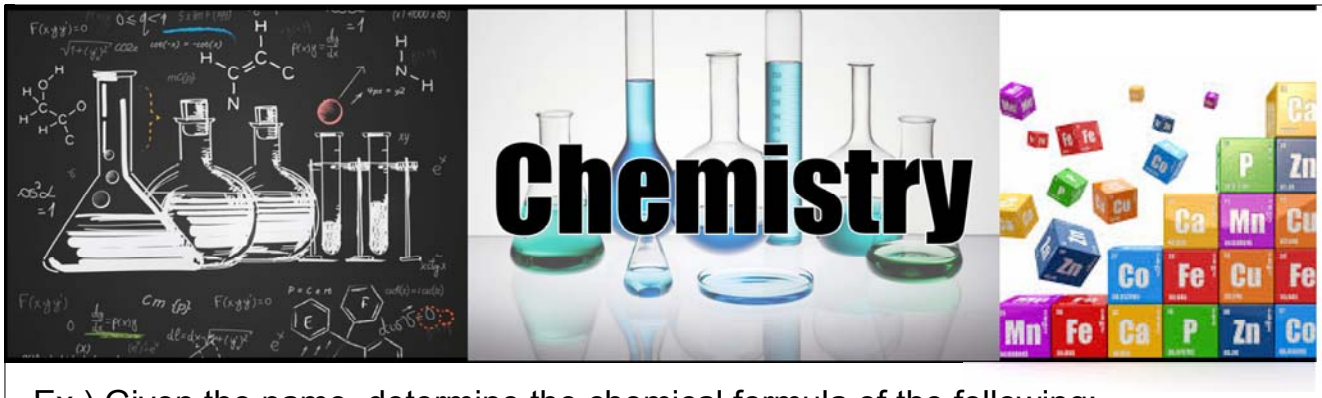

Ex.) Given the name, determine the chemical formula of the following:

a) iron (II) chloride (2+) (1−) $FeCl_{2(s)}$

b) chromium (II) sulfide 2+ 2− $CrS_{(s)}$

c) manganese (IV) fluoride 4+ − $MnF_{4(s)}$

Worksheet