

Lesson 8 Solubility Table Key - questions came from lesson 6

Give the formula for each.

16. ⁺ potassium ²⁻ sulfate $K_2SO_4(aq)$
17. ²⁺ zinc ⁻ bromide $ZnBr_2(aq)$
18. ³⁺ aluminium ⁻ hypochlorite $Al(ClO)_3(aq)$
19. ²⁺ copper (II) ²⁻ carbonate $CuCO_3(s)$
20. ²⁺ copper (II) ⁻ hydrogen sulfide $Cu(HS)_2(s)$
21. ⁺ silver ²⁻ dichromate $Ag_2Cr_2O_7(s)$
22. ³⁺ ruthenium (III) ²⁻ sulfide $Ru_2S_3(s)$
23. ²⁺ calcium ⁻ hydroxide $Ca(OH)_2(s)$
24. ²⁺ mercury (II) ⁻ fluoride $HgF_2(aq)$
25. ⁴⁺ platinum (IV) ⁻ chloride $PtCl_4(aq)$
26. ²⁺ polonium (II) ²⁻ sulfite $PoSO_3(s)$
27. ³⁺ titanium (III) ²⁻ sulfate $Ti_2(SO_4)_3(aq)$
28. ²⁺ zinc ³⁻ phosphate $Zn_3(PO_4)_2(s)$
29. ³⁺ gallium ⁻ hydroxide $Ga(OH)_3(s)$
30. ⁺ potassium ²⁻ chromate $K_2CrO_4(s)$
31. ³⁺ aluminium ²⁻ sulfite $Al_2(SO_3)_3(s)$
32. ²⁺ zinc ⁻ hydrogen sulfide $Zn(HS)_2(s)$

33. ³⁺ aluminium ³⁻ borate $AlBO_3(s)$
34. ²⁺ copper (II) ²⁻ carbonate $CuCO_3(s)$
35. ⁺ ammonium ²⁻ dichromate $(NH_4)_2Cr_2O_7(s)$